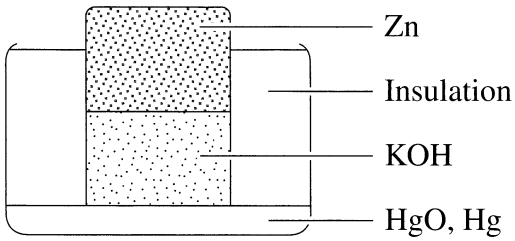
Question 27 (2 marks)

The diagram shows a particular cell with relevant half equations. **2**



$$Zn(s) + 2OH^-(aq) \rightarrow ZnO(s) + H_2O(l) + 2e^-$$

$$HgO(s) + H_2O(l) + 2e^- \rightarrow Hg(l) + 2OH^-(aq)$$

Identify the anode, cathode and electrolyte for this cell.

~~zinc~~ anode = zinc

cathode = HgO

electrolyte = KOH