Question 26 (4 marks)

A gas is produced when 10.0 g of zinc is placed in 0.50 L of 0.20 mol L^{-1} nitric acid.   **4**

Calculate the volume of gas produced at 25°C and 100 kPa. Include a balanced chemical equation in your answer.

nitric acid: $n = cv$ $Zn + H_2 NO_2 \longrightarrow$

$\qquad n = 0.2 \times 0.5$

$\qquad = 0.1 \, mol$